

## CI-340 Equations

**1a.** W: Mass flow rate per leaf area (mol/m<sup>2</sup>/s) for an **open system**.

$$W = \frac{V}{60} \times \frac{273.15}{T_a} \times \frac{P}{1.013} \times \frac{1}{22.41} \times \frac{10000}{A}$$

$$W = 2005.39 \times \frac{V \times P}{T_a \times A}$$

Where *V*: volume flow rate (liters/minute)

*T<sub>a</sub>*: air temperature (K)

*P*: atmospheric pressure (bar)

*A*: leaf area (cm<sup>2</sup>)

60: converts minutes into seconds

22.41: the volume of one mole of any gas at a standard temperature of 273.15K and the standard pressure of 1.013 bar (liters/mol).

10000: converts cm<sup>2</sup> into m<sup>2</sup>

**1b.** W: Mass flow rate per leaf area (mol/m<sup>2</sup>/s) for a **closed system**

$$W = \frac{V}{\Delta t} \times \frac{273.15}{T_a} \times \frac{P}{1.013} \times \frac{1}{22.41} \times \frac{10000}{A}$$

$$W = 2005.39 \times \frac{V \times P}{\Delta t \times T_a \times A}$$

Where *V*: leaf chamber volume (liters)

*T<sub>a</sub>*: air temperature (K)

*P*: atmospheric pressure (bar)

*A*: leaf area (cm<sup>2</sup>)

$\Delta t$ : time interval (seconds)

22.41: the volume of one mole of any gas at a standard temperature of 273.15K and the standard pressure of 1.013 bar (liters/mol).

10000: converts cm<sup>2</sup> into m<sup>2</sup>